General Chemistry 2 Quarter 1- Module 2 - Types of Intermolecular Forces

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**Pre-test**

1. D 2. B 3. A 4. C 5. A

**Elicit**

1. Atom 2. Molecule 3. Cation 4. Anion 5. Polar 6. Non-Polar

**Engage** (Questions):

1. It gets dissolves in to the water, then stirring it lives it looking like any ordinary water
2. Because salt is polar molecule and water is good solvent
3. Yes, because the water molecules pull the sodium and chloride ions apart, breaking the ionic bond that held them together.
4. From the bonds between the water molecule and salt and from the attraction of hydrogen atoms of the water to the chloride atoms of salt causing it’s bonds to split apart.

**Elaborate**

|  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- |
| Substance | Polar/Non-Polar | London Dispersion Forces | Dipole-Dipole Forces | Hydrogen Bond | Ion-Dipole Forces | Ionic Bond |
|  | Polar |  |  |  |  |  |
|  | Non-Polar  Polar |  |  |  |  |  |
|  | Polar  Polar |  |  |  |  |  |

**EVALUATE**

1. 1. Intermolecular Forces
2. Weakest
3. Atom/Molecule
4. Atom/Molecule
5. Temporary dipole
6. Nonpolar
7. Dipole-Dipole
8. Positive/Negative
9. Positive/Negative
10. Hydrogen Bond
11. Oxygen/Fluorine
12. Oxygen/Fluorine
13. Cation/Anion
14. Cation/Anion
15. Solutions

**POST-TEST**

1. C

2. D

3. A

4. B

5. A